Name _____ Class _____ Date _____

Activity

Chemical Reactions

CHAPTER 2

Balancing Equations

1. $Na + O_2 \rightarrow Na_2O$

$\boxed{4}\ Na + O_2 \rightarrow \boxed{2}\ Na_2O$

2. $H_2 + O_2 \rightarrow H_2O$

$\boxed{2}\ H_2 + O_2 \rightarrow \boxed{2}\ H_2O$

3. $Na_2SO_4 + CaCl_2 \rightarrow CaSO_4 + NaCl$

$Na_2SO_4 + CaCl_2 \rightarrow CaSO_4 + \boxed{2}\ NaCl$

4. $Al_2O_3 \rightarrow Al + O_2$

$\boxed{2}\ Al_2O_3 \rightarrow \boxed{4}\ Al + \boxed{3}\ O_2$

5. $N_2 + H_2 \rightarrow NH_3$

$N_2 + \boxed{3}\ H_2 \rightarrow \boxed{2}\ NH_3$

6. $Fe + H_2O \rightarrow Fe_3O_4 + H_2$

$\boxed{3}\ Fe + \boxed{4}\ H_2O \rightarrow Fe_3O_4 + \boxed{4}\ H_2$

7. $P_4 + O_2 \rightarrow P_4O_{10}$

$P_4 + \boxed{5}\ O_2 \rightarrow P_4O_{10}$

8. $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$

$\boxed{2}\ C_6H_6 + \boxed{7}\ O_2 \rightarrow \boxed{4}\ CO_2 \quad \boxed{6}\ H_2O$

9. $SiCl_4 \rightarrow Si + Cl_2$

$SiCl_4 \rightarrow Si + \boxed{2}\ Cl_2$

10. $C + H_2 \rightarrow CH_4$

$C + \boxed{2}\ H_2 \rightarrow CH_4$

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